

IUPAC Task Group on Atmospheric Chemical Kinetic Data Evaluation – Data Sheet III.A2.53 iClOx38

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$$\Delta H^\circ = 43.2 \text{ kJ}\cdot\text{mol}^{-1}$$

Low-pressure rate coefficients

No direct measurements are available.

Preferred Values

$$k_0 = 1.6 \times 10^{-19} \exp(-3810/T) [\text{N}_2] \text{ s}^{-1} \text{ over the temperature range 200-300 K.}$$

Reliability

$$\Delta \log k_0 = \pm 0.5 \text{ at 240 K.}$$

$$\Delta E/R = \pm 500 \text{ K.}$$

Comments on Preferred Values

The preferred values are from the preferred values for the reverse reaction $k_0 = 6.2 \times 10^{-32} (T/300)^{-4.7} [\text{N}_2] \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$, which were based on the measurements by Parr et al. (1990) and Burkholder et al. (1993), and the equilibrium constant $K_C = 9.7 \times 10^{-24} \exp(4940/T) \text{ cm}^3 \text{ molecule}^{-1}$ from Burkholder et al. (1993) and Hayman and Cox (1989).

High-pressure rate coefficients Rate coefficient data

No direct measurements are available.

Preferred Values

$$k_\infty = 2.5 \times 10^{12} \exp(-4940/T) [\text{N}_2] \text{ s}^{-1} \text{ over the temperature range 200-300 K.}$$

Reliability

$$\Delta \log k_\infty = \pm 0.5 \text{ at 240 K.}$$

$$\Delta E/R = \pm 500 \text{ K.}$$

Comments on Preferred Values

See comments on Preferred Values for k_0 . Based on k_∞ for the reverse reaction of $k_\infty = 2.4 \times 10^{-11} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$. Falloff curves constructed with $F_c = 0.6$.

The following text-line combines the preferred values for the high and low pressure limiting rate coefficients to generate a single, cut-and-paste expression for calculation of k :

$$= \frac{(1.6e-19 \cdot \exp(-3810/T)) \cdot M \cdot (2.5e12 \cdot \exp(-4940/T))}{(1.6e-19 \cdot \exp(-3810/T)) \cdot M + (2.5e12 \cdot \exp(-4940/T))} \cdot 10^{\log_{10}(0.6)} \cdot \left(1 + \frac{\log_{10}((1.6e-19 \cdot \exp(-3810/T)) \cdot M / (2.5e12 \cdot \exp(-4940/T)))}{(0.75 - 1.27 \cdot \log_{10}(0.6))} \right)^2$$

The molecular density, $M = 7.243 \times 10^{21} P(\text{bar})/T(\text{K})$

References

- Burkholder, J. B., Mauldin III, R. L., Yokelson, R. J., Solomon, S. and Ravishankara, A. R.: J. Phys. Chem., 97, 7597, 1993.
- Hayman, G. D. and Cox, R.A.: Chem. Phys. Lett., 155, 1, 1989.
- Parr, A. D., Wayne, R. P., Hayman, G. D., Jenkin, M. E. and Cox, R. A.: Geophys. Res. Lett., 17, 2357, 1990.