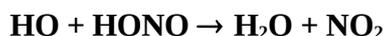


IUPAC Task Group on Atmospheric Chemical Kinetic Data Evaluation – Data Sheet NO_x9

Website: <http://iupac.pole-ether.fr>. See website for latest evaluated data. Data sheets can be downloaded for personal use only and must not be retransmitted or disseminated either electronically or in hardcopy without explicit written permission.

This data sheet last evaluated: 28th July 2007; no revision of preferred values.



$$\Delta H^\circ = -166.3 \text{ kJ mol}^{-1}$$

Rate coefficient data

$k/\text{cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$	Temp./K	Reference	Technique/ Comments
<i>Absolute Rate Coefficients</i>			
$1.80 \times 10^{-11} \exp[-(390 \pm 80)/T]$	278-342	Jenkin and Cox, 1987	MM-RA
$(4.9 \pm 0.5) \times 10^{-12}$	298		
$2.8 \times 10^{-12} \exp[(260 \pm 140)/T]$	298-373	Burkholder et al., 1992	PLP-LIF (a)
$(7.05 \pm 1.41) \times 10^{-12}$	298		
<i>Relative Rate Coefficients</i>			
$(6.0 \pm 0.4) \times 10^{-12}$	296	Cox et al., 1976	(b)

Comments

- (a) HO radicals were monitored by LIF with HONO in excess. [HONO] and [NO₂] were also monitored using *in situ* diode-array spectroscopy.
- (b) Photolysis of HONO-NO and HONO-H₂ mixtures at a total pressure of 1 bar of N₂ or N₂-O₂ diluent. A value of $k/k(\text{HO} + \text{H}_2) = 945 \pm 48$ was obtained. The value of k given in the table is calculated using $k(\text{HO} + \text{H}_2) = 6.4 \times 10^{-15} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$ at 296 K (IUPAC, 2007).

Preferred Values

$$k = 6.0 \times 10^{-12} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1} \text{ at } 298 \text{ K.}$$

$$k = 2.5 \times 10^{-12} \exp(260/T) \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1} \text{ over the temperature range } 290\text{-}380 \text{ K.}$$

Reliability

$$\Delta \log k = \pm 0.15 \text{ at } 298 \text{ K.}$$

$$\Delta(E/R) = \pm 260 \text{ K.}$$

Comments on Preferred Values

There are significant differences between the two direct studies of Jenkin and Cox (1987) and Burkholder et al. (1992), particularly with respect to the temperature dependence of k . The work of Burkholder et al. (1992) has substantially better precision than that of Jenkin and Cox (1987). Furthermore, the small negative temperature dependence of k found by Burkholder et al. (1992) is consistent with that observed for the analogous reaction of HO radicals with HONO₂.

The preferred value of k at 298 K is an average of the room temperature rate coefficients of Jenkin and Cox (1987), Burkholder et al. (1992) and Cox et al. (1976). The temperature dependence of k is that

given by Burkholder et al. (1992), with the pre-exponential factor being adjusted to fit the preferred value of k at 298 K.

References

- Burkholder, J. B., Mellouki, A., Talukdar, R. and Ravishankara, A. R.: *Int. J. Chem. Kinet.* 24, 711, 1992.
Cox, R. A., Derwent, R. G. and Holt, P. M.: *J. Chem. Soc. Faraday Trans. 1*, 72, 2031, 1976.
IUPAC: <http://iupac.pole-ether.fr>, 2007.
Jenkin, M. E. and Cox, R. A.: *Chem. Phys. Lett.* 137, 548, 1987.