

IUPAC Task Group on Atmospheric Chemical Kinetic Data Evaluation – Data Sheet CGI_8

Website: <http://iupac.pole-ether.fr> See website for latest evaluated data. Datasheets can be downloaded for personal use only and must not be retransmitted or disseminated either electronically or in hardcopy without explicit written permission.

This datasheet updated: May 2015; .

CH₂OO + CH₃(CO)CH₃ → products

Rate coefficient data

| $k/\text{cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$ | Temp /K | Reference | Technique/Comments |
|--|------------|----------------------|--------------------|
| Absolute Rate Coefficients | | | |
| $(2.3 \pm 0.3) \times 10^{-13}$ | 293 | Taatjes et al., 2012 | PLP-PIMS(a) |

Comments

(a) CH₂OO (formaldehyde oxide) was produced by the reaction of CH₂I + O₂. CH₂I was generated by 248-nm laser photolysis of diiodomethane, CH₂I₂, at 293 K and 4 Torr total pressure in a large excess of O₂. The reacting mixture was monitored by tunable synchrotron photoionization mass spectrometry, which allowed characterisation of the PIMS for CH₂OO and its reaction products over the region 9.5 – 11.5 eV, and time-resolved direct detection of CH₂OO at $m/z = 46$ amu. The first order decay CH₂OO in the presence of excess known concentrations of acetone was used to determine the rate constants. The uncertainty limits are 95%, based on unweighted linear fit of [CH₃(CO)CH₃] dependence of decay lifetimes. Secondary ozonide (3,3-dimethyl-1,2,4-trioxalane) was identified as a reaction product from its PIMS aided by quantum chemical calculations.

Preferred Values

| Parameter | Value | T/K |
|--|-----------------------|-----|
| $k/\text{cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$ | 2.3×10^{-13} | 298 |
| <i>Reliability</i> | | |
| $\Delta \log k$ | ± 0.3 | 298 |

Comments on Preferred Values

The direct photoionization mass spectrometric detection of formaldehyde oxide (CH₂OO) as a product of the photolysis of CH₂I₂ in the presence of O₂ (Welz et al., 2012) has allowed detailed investigation of the kinetics of the Criegee intermediates, including the reaction: CH₂OO + CH₃(CO)CH₃ → products. The rate measurements appear to be precise and consistent with the emerging reactivity pattern for CH₂OO reactivity. The higher uncertainty reflects the single investigation.

References

Taatjes, C.A., Welz, O., Eskola, A.K., Savee, J. D., Osborn, D. L., Lee, E.P.F., Dyke, J.M., Mok, D.W.K., Shallcross, D. E., and Percival, C. J., *Phys Chem.Chem.Phys.*, 14, 10391, 2012.