IUPAC Task Group on Atmospheric Chemical Kinetic Data Evaluation - Data Sheet Het\_Org4

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This data sheet evaluated: June 2016; last change in preferred values: June 2016.

**NO3 + unsaturated acids**

**Experimental data**

|  |  |  |  |
| --- | --- | --- | --- |
| *Parameter* | Temp./K | Reference | Technique/ Comments |
| γ0 |  |  |  |
| (7.9 ± 1.2) × 10-3 (linoleic acid, conj.) | 273 | Moise et al., 2002 | CWFT-AS (a) |
| (15.0 ± 2.0) × 10-3 (linoleic acid, unconj.) | 288 |  |  |
| > 0.12(linoleic acid, conj.) | 278-298 | Gross et al., 2009 | CWFT-CIMS (b) |
| > 0.13(linoleic acid, unconj.) | 288 |  |  |
| > 0.07(oleic acid) | 285-302 |  |  |
| 1.01 ± 0.18 (linoleic acid, conj.) | 295 | Zhao et al., 2011 | AFT (c) |
| 0.53 ± 0.12 (linoleic acid, unconj.) | 295 |  |  |
| 0.27 ± 0.06 (oleic acid, conj.) | 295 |  |  |
| 0.72 ± 0.17 (linolenic acid, conj.) | 295 |  |  |
| (2-6) × 10-3 (abietic) | 298 | Knopf et al., 2011 | CWFT-CIMS (d) |
| (linoleic) | 298 |  |  |
| (oleic) | 298 |  |  |
| 0.14 ± 0.03 (oleic) | RT | Renbaum-Wolff and Smith, (2012) | AFT (e) |
|  |  |  |  |

Abietic acid (C20H30O2) is Abieta-7,13-dien-18-oic acid, Linoleic acid is C18H32O2 ((9*Z*,12*Z*)-9,12-octadecadienoic acid), linoleic acid (conj.) is a mixture of cis- and trans-9,11 octadecadienoic acid and -10,12 octadecadienoic acid, linolenic acid is C18H30O2 ((9*Z*,12*Z*,15*Z*)-9,12,15-octadecatrienoic acid), oleic acid is C18H34O2, ((9*Z*)-Octadec-9-enoic acid), RT = room temperature.

Comments

1. Rotating, coated wall flow tube. NO3 (5-50 × 1011 molecule cm‑3) was formed by the thermal dissociation of N2O5 and detected by diode laser absorption at 662 nm.
2. Rotating, coated wall flow tube. NO3 (3.5-16 × 1010 molecule cm‑3) was formed by the thermal dissociation of N2O5 and detected following ionisation by I-. The uptake coefficient for diethyl sebacate decreased to ~0.06 ± 0.015 upon freezing at 268-283 K.
3. Mixed-phase, relative rates method. Loss of particle-phase organic monitored relative to loss of gas-phase reference reactant (isoprene, 2,3,-dimethyl-1,3-butadiene, cyclopentane and 2-methyl-2-2butene).
4. NO3 (4-300 × 109 molecule cm‑3) was formed by the thermal dissociation of N2O5 and detected following ionisation by I-. The larger values of γ0 were obtained when using low NO3 concentrations.
5. Mixed-phase, relative rates method. NO3 generated from 355 nm photolysis of a mixture of HONO / HNO3. Particle-phase oleic acid and gas-phase isoprene (reference reactant) monitored by Aerosol CIMS.

**Preferred Values**

|  |  |  |
| --- | --- | --- |
| **Parameter** | **Value** | **T/K** |
| *α*b | 1 |  |
| *k*b (M-1 s-1) | 3 × 108 | 280-300 |

*Reliability*

|  |  |  |
| --- | --- | --- |
|  |  |  |
| Δ log(*k*b) | 0.5 |  |
|  |  |  |

Comments on Preferred Values

There is considerable scatter in the uptake coefficients obtained for different unsaturated acids and also when comparing results of different groups for the same acid (e.g. linoleic). Good agreement between expeirments using bulk samples and particlulate linoleic and oleic acids suggest that the uptake coefficient is > 0.2.

If the reaction between NO3 and the unsaturated acid (of concentration [HC], in units of mol L-1) takes place throughout the particle, the uptake coefficient can be described by:

Where *k*b(i) is the liquid-phase rate coefficient for reaction of NO3 with organic species (i)with concentration [HC], *Dl* its diffusion coefficient through the organic matrix and *H* its solubility.

A rough estimate for a generic uptake coefficient for NO3 uptake to unsaturated acids can be made using *kb*= 3 × 108 M-1 s-1 (equivalent to a gas-phase rate constant of 5 × 10-13 cm3 molecule-1 s‑1), *Dl*, = 2 × 10-5 cm2 s-1, and *H* = 0.8 Matm-1, this expression results in a value of γ = 0.3, which is consistent with the experimental data if we assign an uncertainty of a factor 4. The large uncertainty associated with use of a generic rate constant, *k*b, and also use of potentially inappropriate values of *H* and *Dl*  is taken into account by assigning a large uncertainty to *kb*.

Uptake to multicomponent organic mixtures can be approximated by summing the product *k*b(i)[HC](i) and using an average value for *H* and *Dl*. For unreactive or very small particles a correction for the diffuso-reactive length may be important (see guide to datasheets), whereas for very reactive particles, the reaction my be limited to the surface layers of the sample. This may result in uptake coefficients that decrease with exposure time if mixing in the particle is hindered by high viscosity. The NO3 concentration- and time-dependence of the uptake coefficient to a sample of abeitic acid (Knopf et al., 2011) could be described with a multilayer model of surface and bulk reactions and bulk diffusion of NO3 (Shiraiwa et al., 2012).

The reaction of NO3 with the unsaturated organics considered here proceeds largely through addition. Products from the reaction of NO3 with oleic acid are carboxylic acids including hydroxy nitrates, carbonyl nitrates, dinitrates and hydroxydinitrates (Hung et al., 2005; Docherty and Ziemann, 2006; Zhao et al., 2011).

References

Docherty, K. S., and Ziemann, P. J., J. Phys. Chem. A, 110, 3567-3577, 2006.

Gross, S., Iannone, R., Xiao, S., and Bertram, A. K., Phys. Chem. Chem. Phys., 11, 7792-7803, 2009.

Hung, H. M., Katrib, Y., and Martin, S. T., J. Phys. Chem. A, 109, 4517-4530, 2005.

Knopf, D. A., Forrester, S. M., and Slade, J. H., Phys. Chem. Chem. Phys., 13, 21050-21062, 2011.

Moise, T., Talukdar, R. K., Frost, G. J., Fox, R. W. and Rudich, Y.: J. Geophys. Res.-Atmos. 107, D2, doi: 10.1029/2001JD000334,2002.

Renbaum-Wolff, L., Grayson, J. W., Bateman, A. P., Kuwata, M., Sellier, M., Murray, B. J., Shilling, J. E., Martin, S. T., and Bertram, A. K., Proc. Natl. Acad. Sci. U. S. A., 110, 8014-8019, 2013.

Shiraiwa, M., Poschl, U., and Knopf, D. A., Env. Sci. Tech., 46, 6630-6636, 2012.

Zhao, Z. J., Husainy, S., Stoudemayer, C. T., and Smith, G. D., Phys. Chem. Chem. Phys., 13, 17809-17817, 2011.